<table>
<thead>
<tr>
<th>Quantum Number</th>
<th>Name</th>
<th>Explanation</th>
</tr>
</thead>
<tbody>
<tr>
<td>n</td>
<td>principal</td>
<td>electron energy level or shell number. 1, 2, 3,...</td>
</tr>
<tr>
<td>l</td>
<td>subshell</td>
<td>orbital type s, p, d, f. 0, 1, 2, 3,...</td>
</tr>
<tr>
<td>m_l</td>
<td>magnetic</td>
<td>orbital sub-type. integer between and including -l, l</td>
</tr>
<tr>
<td>m_s</td>
<td>spin</td>
<td>0, 1/2</td>
</tr>
</tbody>
</table>

Notes:
- For n = 1, only s orbitals are allowed.
- Small letters of atomic orbitals:
  - s: spherical
  - p: dumbbell
  - d: hourglass
  - f: complex

Spin in main single sub-energy level must be paired.
expectation values. The Schrödinger equation is then derived by using these results and demanding both Galilean invariance of the probability density and Newtonian energy-momentum relations. The general properties of the Schrödinger equation are used to introduce the idea of the wavefunction. The book also includes a discussion of the relationship between the Schrödinger equation and the concept of probability, as well as the role of the wavefunction in quantum mechanics. The text is written in a clear and concise manner, with numerous examples and exercises included to help readers understand the key concepts. A Concise Introduction to Quantum Mechanics is an excellent resource for students and researchers interested in learning about quantum mechanics, as well as for anyone looking to gain a deeper understanding of the subject. The book is highly recommended for anyone interested in the field of quantum mechanics.
null
How To Do Quantum Numbers Chemistry

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